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Answer Key Chapter 13 - Chemistry 2e | OpenStax. 1. The reaction can proceed in both the forward and reverse directions. 3. When a system has reached equilibrium, no further changes in the reactant and product concentrations occur; the forward and reverse reactions continue to proceed, but at equal rates. 5.

### Answer Key Chapter 13 - Chemistry 2e | OpenStax

Chapter 13: Organic Chemistry Naming Practice. 2,4-dichloro-3-methylpentane... 5-carbon.... 5-tert-butylnonane... 9-carbon chain - no.... 4-ethyloctane... 8-carbon chain - octane.... 3-bromo-2-methylpentane... 5-carbon chain.... Name this molecule. Name this molecule.

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Equilibrium cannot be established between the liquid and the gas phase if the top is removed from the bottle because the system is not closed; one of the components of the equilibrium, the Br<sub>2</sub> vapor, would escape from the bottle until all liquid disappeared. Thus, more liquid would evaporate than can condense back from the gas phase to the liquid phase.

### **Answer Key Chapter 13 - Chemistry: Atoms First 2e | OpenStax**

The amount of energy required to spin-flip a nucleus depends both on the strength of the external magnetic field and on the nucleus. At a field strength of 4.7 T, rf energy of 200 MHz is required to bring a <sup>1</sup>H nucleus into resonance, but energy of only 187 MHz will bring a <sup>19</sup>F nucleus into resonance. Calculate the amount of energy required to spin-flip a <sup>19</sup>F nucleus.

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### **Solved: For the following reactions and data, provide the ...**

Chapter Outline. 13.1 Chemical Equilibria. 13.2 Equilibrium Constants. 13.3 Shifting Equilibria: Le Châtelier's Principle. 13.4 Equilibrium Calculations. Imagine a beach populated with sunbathers and swimmers. As those basking in the sun get too hot, they enter the surf to swim and cool off. As the swimmers tire, they return to the beach to rest. If the rate at which sunbathers enter the surf were to equal the rate at which swimmers return to the sand, then the numbers (though not the ...

### **Ch. 13 Introduction - Chemistry 2e | OpenStax**

This is a case in which the solution contains a mixture of acids of different ionization strengths. In solution, the HCO<sub>2</sub>H exists primarily as HCO<sub>2</sub>H molecules because the ionization of the weak acid is suppressed by the strong acid. Therefore, the HCO<sub>2</sub>H contributes a negligible amount of hydronium ions to the solution. The stronger acid, HCl, is the dominant producer of hydronium ions

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